

Research Article

Ammonia-Nitrogen Adsorption by Melon Rind

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Abstract

The presence of NH₃-N in a watery arrangement is a major issue of environment, NH₃-N is the significant water contamination. In this investigation, melon Rind for low-cost adsorbent instead of expensive technologies for removal of NH₃-N from water arrangement, hence to increase the attractiveness of melon rind for absorption of NH₃-N, Rind was treated with NaOH (0.3 M). To detect groups on the surface of the Rind were utilizing the Fourier Transform Infra-red (FTIR). SEM scanning electron microscopy was utilized morphological properties of the pre-arranged Melon Rind, Energy-dispersive X-beam spectroscopy (EDS) was utilized to research the

Rind Surface Elements affecting adsorption are contact time, adsorbent dose, pH, NH₃-N concentration, were investigated. The adsorption limit of NH₃-N diminished as the adsorbent portion expanded. The Kinetic examinations showed best coordinated with the pseudo-first-order and pseudo-second-order models. The Freundlich and Langmuir isotherm models were utilized to depict the NH₃-N adsorption onto the melon Rind adsorbent, the information gotten from the adsorption interaction were best agreeable with the Langmuir model with R²0.9989 and a most q_{max} adsorption limit of 12.161mg/g.

Keywords: Adsorption; NH₃-N; Isotherm; Kinetic; melon rind; EXD; SEM; FTIR

1. Introduction

Water contamination has become a significant overall natural issue annihilating sea-going biological systems, with the improvement of industrialization, different synthetic squanders from mechanical plants are pouring to ecological environments is one of the significant elements in water contamination [1] these squanders incorporate poisonous substances, muck and solvents. Sadly, alkali nitrogen is a typical water impurity in the solvents that affects the climate and represents a danger to human wellbeing [2]. What's more, smelling salts nitrogen is one of the significant nitrogen-containing toxins which may prompts the eutrophication and worsening algal blossoms in environments [3,4]. Nitrogen contamination incorporates a few structures and smelling salts nitrogen (NH₃-N) is the average structure that comes from civil sewage, modern wastewater, landfill leachate and rural non-point source release [6]. Numerous methods are accessible for the expulsion of smelling salts nitrogen from water stage. Compound strategy anaerobic smelling salts expulsion [7] film distillation [8] alkali expulsion from watery arrangement by zeolite [9] of organic product, rinds based adsorbent for eliminating smelling salts nitrogen contaminations from water offers numerous alluring highlights like the exceptional adsorption limit and the way that these materials are minimal expense, non-poisonous and biocompatible [10]. In this examination, a new adsorbent was produced using melon Rind low cost

to eliminate ammonia nitrogen from fluid water arrangement. Rind ware treated with NaOH to improve adsorbent adsorption capacity.

2. Material and Methods

2.1. Preparation of the Adsorbent

The melon Rind was gathered from TYUT organic product market, Shanxi/china. All synthetic substances utilized in this investigation are of standard scientific grade and were acquired from TYUT. The melon Rind were washed to eliminate dust utilizing faucet water, trailed by deionized, dried in a stove at 60°C for 6 days then, at that point cut it in ground into fine Nano-particles utilizing a processor, sieved. Take 10 g melon Rind 100 mL of a 0.3M arrangement of NaOH combination was mixed until uniform and drenched for 24 hours the biomass after treatment was washed with deionized water until the pH of the wash arrangement was close to the nonpartisan reach 7. After this, the adsorbent was dried at 60°C for 24 hours in a drying stove and put away dry until use, to be specific NaOH+ melon Rinds (NMRs).

2.2. Preparation of stock solution

Engineered NH₃-N arrangements were utilized all through the adsorption tests. Alkali chloride salt (NH₄Cl) with determined weight was disintegrated in 1L of refined water for wanted focus. The underlying pH worth of arrangement was changed in accordance with 7.0 utilizing weaken arrangement of 0.1 M of hydrochloric corrosive HCL and sodium hydroxide (NaOH).

2.3. Batch Adsorption test

Batch experiment solutions were engendering by diluting the stock solutions in deionized water. The adsorption investigation of NH₃-N on NMRs by using the following operating parameters: pH (3-9), biosorbent amount (0.1-1 g/L), concentration initial (0.4-10 mg/L), adsorption period (5 min-90min) at 25°C. The effect of pH on adsorption of NH₃-N was investigated using 0.1 g/L biosorbent and initial concentration of 10 mg/L in a 50 mL suspension volume. At optimal pH 7, the effect of biosorbent amount on NH₃-N adsorption was examined by biosorbent increasing from 0.1g/L to 1g/L, at 25°C, the fancy volume of 50 mL was shaken at 180 rpm. aim of the kinetic study was to find how long it takes to reach equilibrium using the real amount of biosorbent concentration initial 0.4 to 10 mg/L were used in equilibrium adsorption tests, Optimal dose of adsorbent of 0.1 g/L, using 0.1M HCl and 0.1M NaOH solutions for pH adjustment. The NMPs were isolated from the medium using filtration after equilibration, measure of NH₃-N in the arrangements were dissected in nurse's region at a frequency of 420 nm by the (PerkinElmer spectrophotometer Spector quant Nova 60). At last, the NH₃-N harmony was determined as displayed in Equation the measure of NH₃-N adsorbed was resolved utilizing the condition underneath eq1:

$$q_e = \frac{C_0 - C_e}{m} V \quad (1)$$

Where C_0 and C_e are NH₃-N focuses (mgL⁻¹) before and after the fact adsorption, V is the volume of adsorbate (L) and m is the heaviness of the adsorbent (g).

2.4. Adsorbent characterizations

The characteristics of NMRs such as surface morphology, surface functional group, chemical characteristics were determined through laboratory diagnostic techniques using scanning electron microscopy (SEM)(EDX), Fourier transform infrared (FTIR)

3. Results

3.1. Characterization of Biosorbent

The obtained results of FTIR spectra showed that the surface of NMRs had the different functional groups. The peaks showed in (Figure 1a) for NMPs are related to O-H groups, carbonyl, C-H, hydroxyl and carboxyl. (Figure 1b) In the SEM micrographs, surface of NMPs showed many pours on the surface of sorbent it suitable for NH₃-N adsorption. (Figure 1c) shows EDX analysis which was performed for NMRs in order to determine the chemical compound of adsorbent. The obtained results indicate that the adsorbent, in addition to main composition, has different amounts of other chemical compounds.

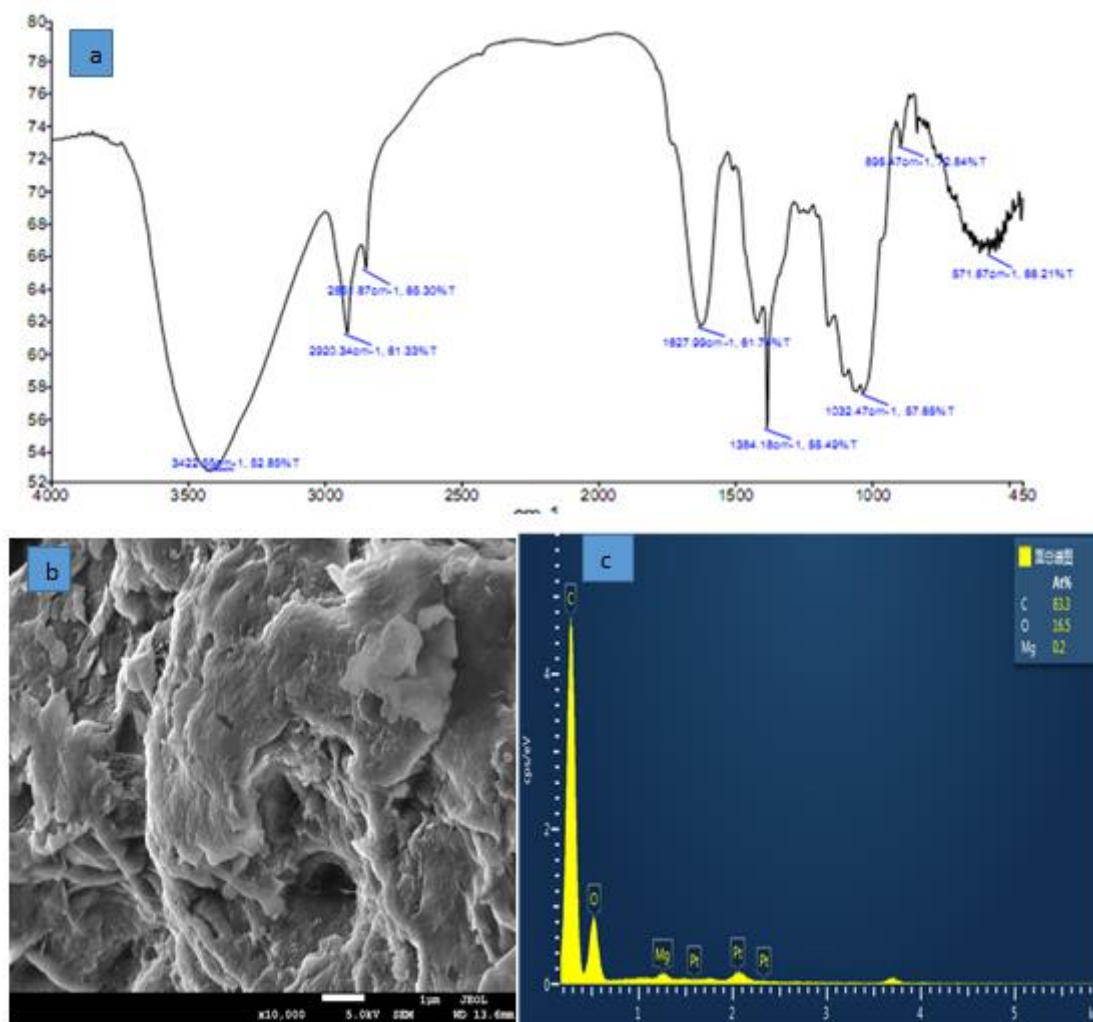


Figure 1: FTIR (a) SEM image (b), EDS spectrum (c)

3.2 Effects of parameters on adsorption process

3.2.1. Effect of initial pH: The underlying pH arrangement is perhaps the most basic controlling boundaries in adsorption measures. The impact of introductory pH on the adsorption cycle was seen by taking the normal pH upsides of water where the worth lies in scope of 3 to 11. The adsorption limit of ammonia nitrogen q_e (mg/g) at various pH is

displayed in figure 2 for the underlying arrangement centralization of 10 mg/L. The adsorption limit of NMPs on the NH_3-N measure shows least at esteem at pH 3 and 11(0.17 and 0.45), expanded to 0.56 mg/g at pH 5 and arrived at greatest adsorption limit (0.745mg/g) at pH 7. The adsorption esteems diminished to 0.665 mg/g as the pH worth of 9.

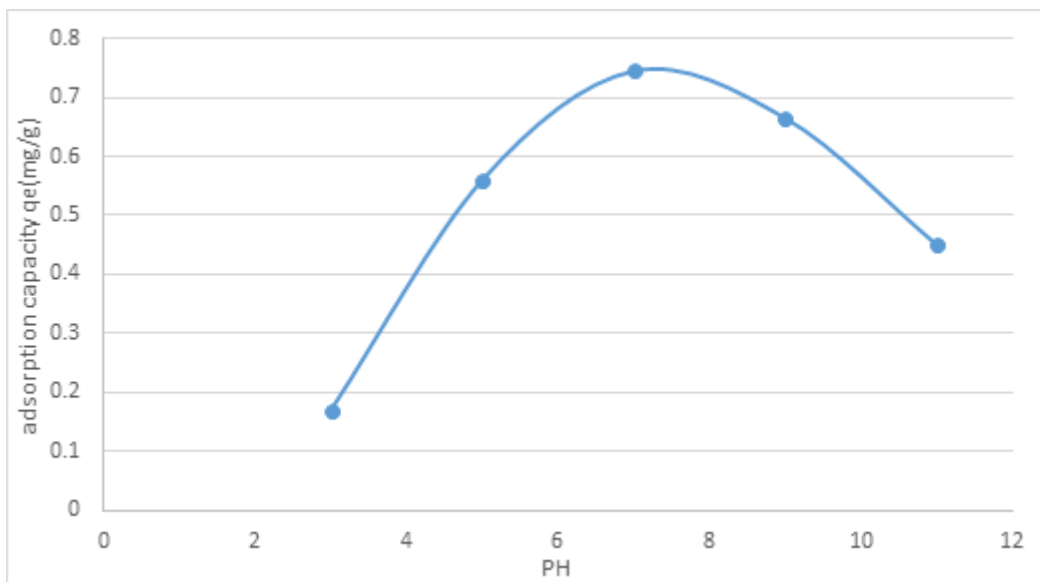


Figure 2: Effect of initial pH on the adsorption capacity of NMPs.

3.2.2. Effect of adsorbent dosage: As shown in figure 3, the adsorption limit of (NMRs) diminished with the expansion of adsorbent dose from 0.1g to 1g. This impact can be credited to a lessening number of

dynamic surface locales for the $\text{NH}_3\text{-N}$ subsequently essentially decreasing the adsorption limit. This was principally because of the covering of adsorbent layers.

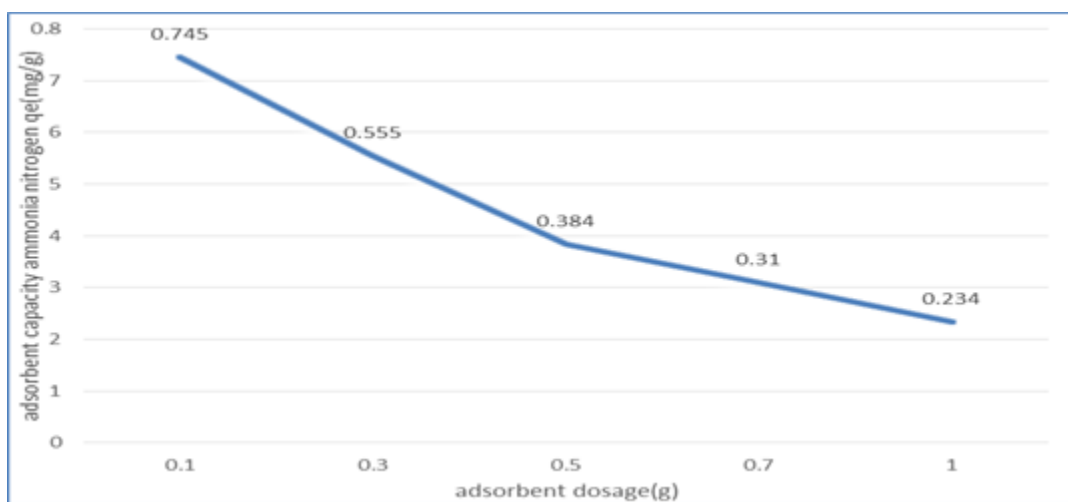


Figure 3: Graph of Adsorption Capacity of (NMRs), q_t (mg/g) against Adsorbent Dosage (g)

3.3.3. Effect of contact time of adsorption: As delineated in figure 4, it tends to be seen that the adsorption limit was quick at the main phase of adsorption (0-20 minutes), then, at that point gradually expanded with time (20-an hour), and getting practically steady following an hour (an hour onwards). This marvel was because of the great accessibility of empty destinations in the primary

phase of adsorption. The adsorption was then trailed by an expansion in shocking powers because of the presence of the adsorbed particles. The adsorption cycle proceeded to a point where the adsorbent couldn't adsorb the encompassing adsorbate because of the absence of empty locales. In this examination, the ideal contact time was an hour.

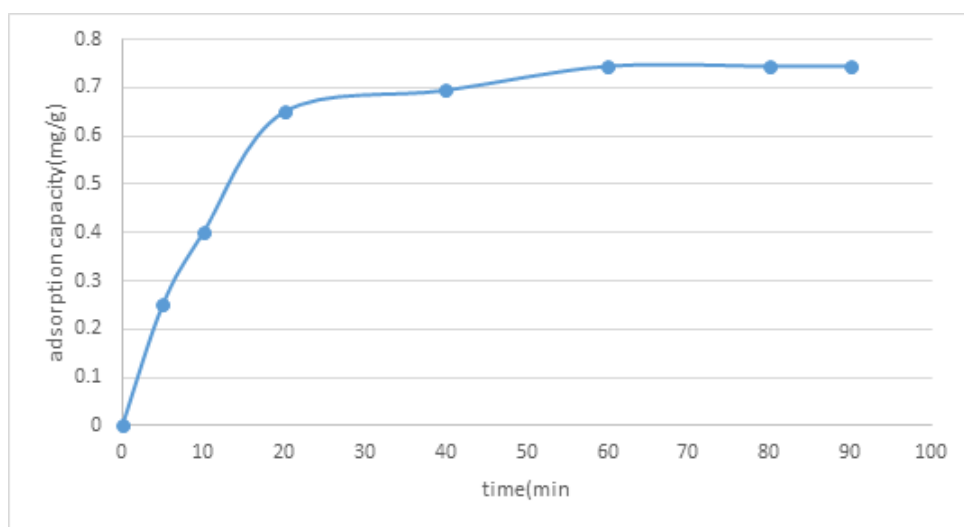


Figure 4: Graph of Adsorption Capacity of (NMRs), q_t (mg/g) against Contact Time (min)

3.2.4. Effect of initial concentration of $\text{NH}_3\text{-N}$ solution: From figure 5, clearly by expanding the underlying convergence of $\text{NH}_3\text{-N}$ arrangement from 0.4 mg/l to 10mg/l, the adsorption limit of the (NMRs), expanded from 0.0283 mg/g to 0.745 mg/g.

It was on the grounds that higher $\text{NH}_3\text{-N}$ fixation gave a more noteworthy main impetus to adsorption as more $\text{NH}_3\text{-N}$ particles can be adsorbed by the (NMRs).

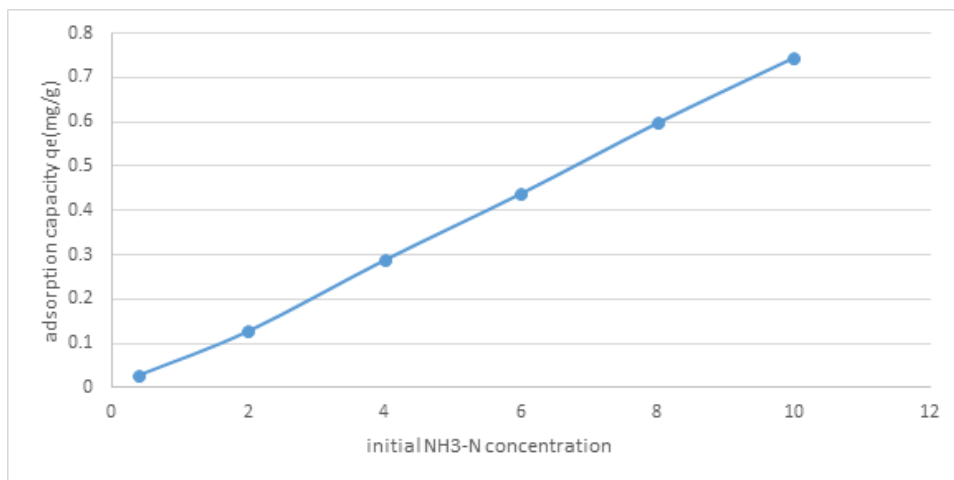


Figure 5: Graph of Adsorption Capacity of (NMRs), qe (mg/g)

3.3. Kinetic study

The adsorption motor was performed by shifted adsorption contact time which was 0, 10, 20, 30, 40, 60, 80, and an hour and a half to continue active fitting and decide adsorption rate. The two diagrams of pseudo-first-request and pseudo-second-request show various practices. The worth of R2 of pseudo

second-request modular was 0.9905 which closer to 1 contrasted with pseudo-first-request, 0.9106. Consequently, pseudo-second-request model was more ideal for adsorption dynamic which the worth of R2 is closer to 1 that show a decent numerical fit (figures 6,7).

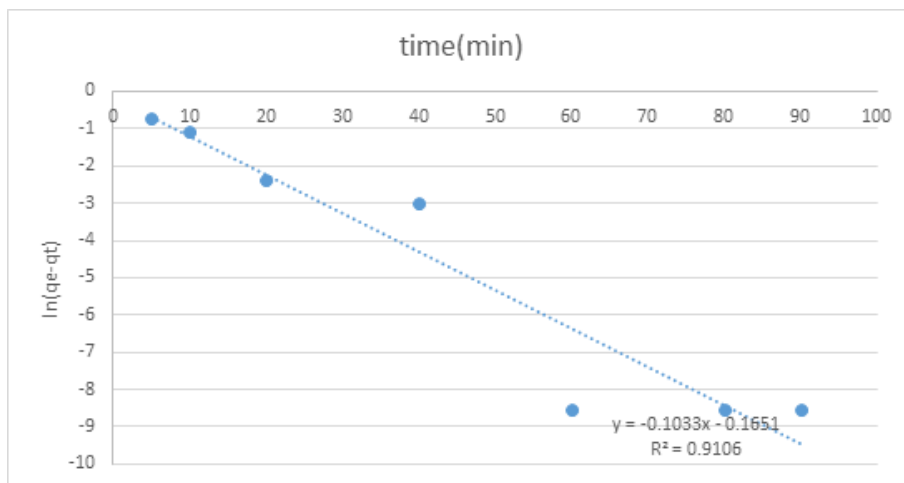


Figure 6: Graph of pseudo-first-order.

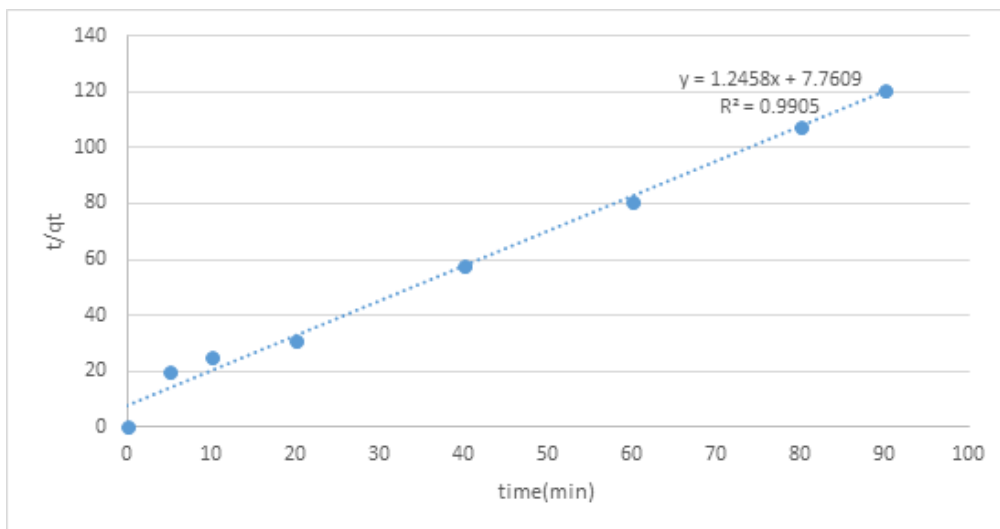


Figure 7: Graph of pseudo second order.

3.4. Adsorption isotherm

The harmony adsorption information was depicted by two of the most much of the time utilized isotherm models: Langmuir and Freundlich isotherm models. The Langmuir isotherm model can be communicated in a linearized condition as eq2:

$$\frac{C_e}{q_e} = \frac{1}{q_m K_L} + \frac{C_e}{q_m} \quad (2)$$

Where C_e is the harmony centralization of NH₃-N (mg/L), q_e is the measure of NH₃-N adsorbed at balance (mg/g), q_m is the greatest adsorption limit (mg/g) and K_L is the Langmuir isotherm consistent

identified with the free energy of adsorption (L/mg). Figure 8 showed the plot of $1/q_e$ against beginning focuses ($1/C_e$) which gave a straight-line diagram. The Langmuir consistent (K_L) and most extreme adsorption limit were acquired from the incline and capture of the chart individually. Table 1 likewise showed a Langmuir plot with a relationship coefficient of 0.9991, which was more prominent than 0.99, in this way, demonstrating that the information acquired for the adsorption interaction fitted well into the Langmuir Isotherm Model.

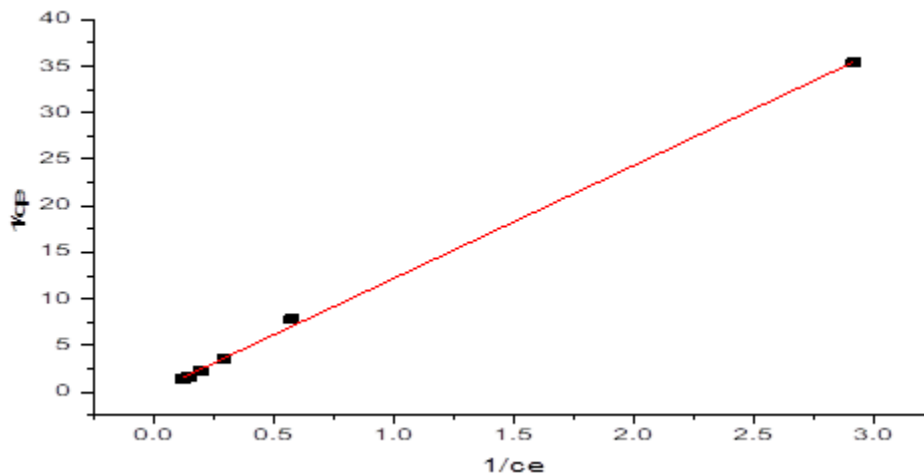


Figure 8: Langmuir Isotherm graph on NH3-N adsorption by NMPs

Freundlich equation can be expressed as: eq3

$$\log q_e = \log K_F + \frac{1}{n} \log C_e \quad (3)$$

Where, K_F is the constant related to adsorbent capacity (mgg^{-1}) and $\frac{1}{n}$ is an empirical parameter related to the sorption intensity of the adsorbent. The plot of $\log q_e$ against $\log C_e$ gave a straight-line graph

as seen in Figure 9, with a slope of $\frac{1}{n}$ and an intercept of K_F . The Freundlich Isotherm model plot of data obtained from the NH3-N adsorption onto NMPs with gave a correlation coefficient of table 2, $R^2 = 0.9971$. The values of K_F and $\frac{1}{n}$ obtained were 0.080707 And 1.02897 respectively.

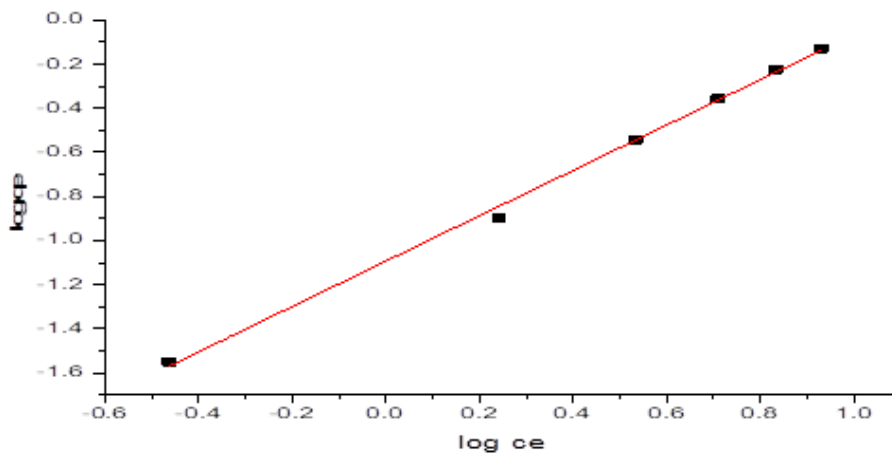


Figure 9: Freundlich Isotherm graph on NH3-N adsorption by NMPs

Intercept	Slope	$q_{\max}(\text{mg/g})$	K_L	R_L	R^2
0.08223	12.14519	12.1610118	0.00677058	0.93658757	0.9989

Table 1: Langmuir Parameters

Intercept	Slope	1/n	K_f	R^2
-1.09309	1.02897	1.02897	0.080707	0.9971

Table 2: Freundlich Parameters

4. Conclusion

This Research was done to investigate biosorbent melon Rinds, as a new biosorbent, for treating $\text{NH}_3\text{-N}$ from watery arrangement. The exploratory outcomes showed that melon Rind treated by 0.3M NaOH (NMRs) can significantly further develop the sorption expulsion for $\text{NH}_3\text{-N}$ Chemicals alteration for melon Rinds and pH showed substantial impact on $\text{NH}_3\text{-N}$ evacuation, while contact time introduced brief impact. The energy study showed that the sorption conduct better fit the pseudo-second-model condition than the Langmuir and Freundlich equation (99). The FTIR retention range investigation of melon Rinds and NMPs prior and then afterward sorption uncovered that the movement deferent gatherings at the outside of the melon Rinds SEM and XRD portrayal tests completed on the (NMRs) to see better adsorbent surface. In light of the adsorbent the most appropriate to be utilized for $\text{NH}_3\text{-N}$ adsorption measure. In this investigation, 0.1 g of (NMRs) adsorbent was demonstrated the most ideal dose for the adsorption of $\text{NH}_3\text{-N}$. Concerning contact time examination, an hour was the most ideal contact time for adsorption of $\text{NH}_3\text{-N}$. 10mg/L of $\text{NH}_3\text{-N}$ arrangement was the most proficient starting focus for adsorption.

Conflicts of Interest

The authors declare that they have no conflicts of interest

Availability of data and materials

All data, models, and code generated or used during the study appear in the submitted article.

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